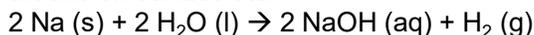


2.4 Periodicity of Period 3

Trends in the reactions of the elements with water, limited to Na and Mg

Learn the equations

Sodium reacts with cold water. It fizzes around on surface etc.



Magnesium reacts very slowly with cold water to form the hydroxide but reacts more readily with **steam** to form the oxide

$$\text{Mg (s)} + \text{H}_2\text{O (g)} \rightarrow \text{MgO (s)} + \text{H}_2 \text{(g)}$$

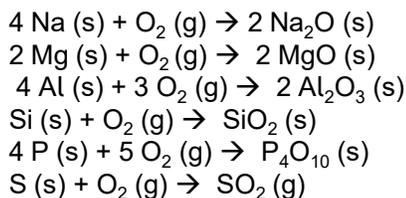
Trends in the reactions of the elements Na, Mg, Al, Si, P and S with oxygen

The elements all react with oxygen to form oxides.

Sodium burns with a **yellow flame** to produce a **white solid**.

Mg, Al, Si and P burn with a **white flame** to give **white solid** smoke.

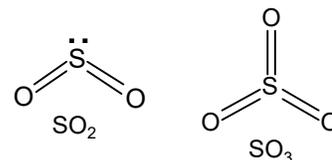
S burns with a **blue flame** to form an acidic **choking gas**.



You should be able to write these equations.

Learn the formulae of the oxides

Sodium is stored under oil and phosphorus under water to stop these elements coming into contact and reacting with air.



A survey of the properties of the oxides of Period 3 elements

Understand the link between the physical properties of the highest oxides of the elements Na → S and their structure and bonding.

Ionic oxides

The metal oxides (Na₂O, MgO, Al₂O₃) are ionic. They have high melting points. They have **ionic giant lattice structures**. The strong electrostatic forces of attraction between oppositely charged ions mean they have high melting points. They are ionic because of the large electronegativity difference between metal and oxygen.

The increased charge on the cation going from Na to Al makes the ionic forces stronger (bigger lattice enthalpies of dissociation) so leading to increasing melting points.

Al₂O₃ is ionic but does show some covalent character. This can be explained by the electronegativity difference being less big or alternatively by the small aluminium ion with a high charge being able to get close to the oxide ion and distorting the oxide charge cloud.

To prove that the above compounds contain ions experimentally - **melt** the solids and show they conduct electricity.

Macromolecular oxides

SiO₂ is macromolecular: It has many very strong covalent bonds between atoms. High energy needed to break the many strong covalent bonds – very high melting + boiling point

Simple molecular oxides:

P₄O₁₀, SO₂ are simple molecular with weak intermolecular forces between molecules (Van der Waals + permanent dipoles) so have lower mp's. They are covalent because of the small electronegativity difference between the non-metal and O atoms.

P₄O₁₀ is a molecule containing 4P's and 10 O's. As it is a bigger molecule and has more electrons than SO₂ it will have larger Van der Waals forces between molecules and a higher melting point.

Aluminium metal is protected from corrosion in moist air by a thin layer of aluminium oxide. The high lattice strength of aluminium oxide and its insolubility in water make this layer impermeable to air and water.

The reactions of the oxides of the elements Na → S with water

Metal ionic oxides tend to react with water to form hydroxides which are alkaline.

$\text{Na}_2\text{O (s)} + \text{H}_2\text{O (l)} \rightarrow 2\text{Na}^+ \text{(aq)} + 2\text{OH}^- \text{(aq)}$ **pH 13 (A violent exothermic reaction that produces a colourless solution)**

The ionic oxides are basic because the oxide ions accept protons to become hydroxide ions in this reaction (acting as a Bronsted-Lowry base)

$\text{MgO (s)} + \text{H}_2\text{O (l)} \rightarrow \text{Mg(OH)}_2 \text{(s)}$ **pH 9**

Mg(OH)_2 is only slightly soluble in water as its lattice is stronger so there are fewer free OH^- ions meaning the solution has a lower pH than NaOH.

Know the change in pH of the resulting solutions across the period.

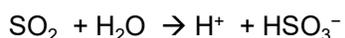
MgO is better than NaOH for treating excess acid in rivers and the stomach as it is only sparingly soluble and weakly alkaline so using an excess would not make the water excessively alkaline.

Al_2O_3 and SiO_2 **do not dissolve** in water because of the high strength of the Al_2O_3 ionic lattice and the SiO_2 macromolecular structure, so they give a neutral **pH 7**

The non-metal, **simple molecular**, covalent, oxides react with water to give acids.

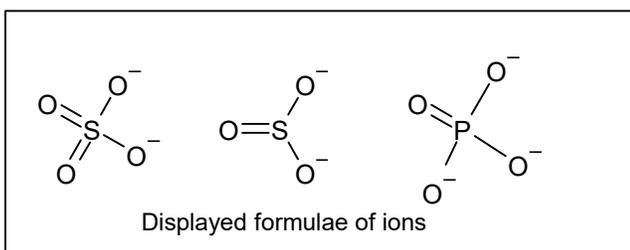
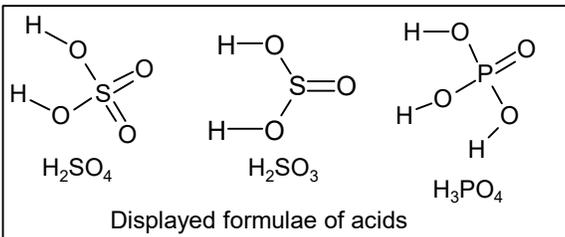
$\text{P}_4\text{O}_{10} \text{(s)} + 6 \text{H}_2\text{O (l)} \rightarrow 4 \text{H}_3\text{PO}_4 \text{(aq)}$ **pH 0 (A violent exothermic reaction that produces a colourless solution)**

$\text{SO}_2 \text{(g)} + \text{H}_2\text{O (l)} \rightarrow \text{H}_2\text{SO}_3 \text{(aq)}$ **pH 3 (weak acid)**



Equations showing formation of ions in solution

$\text{SO}_3 \text{(g)} + \text{H}_2\text{O (l)} \rightarrow \text{H}_2\text{SO}_4 \text{(aq)}$ **pH 0**



The trend is the **ionic metal oxides** show **basic** behaviour and the **non-metal covalent** oxides show **acidic** behaviour.

The slightly intermediate nature of the bonding in aluminium oxide is reflected in its **amphoteric** behaviour: it can act as both a base and an acid.

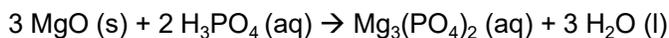
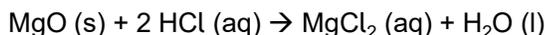
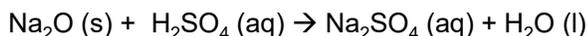
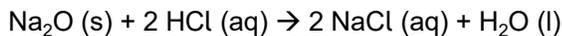
Acid base reactions between period 3 oxides and simple acids and bases.

Rather than learning the equations by rote, learn the pattern.

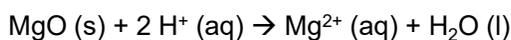
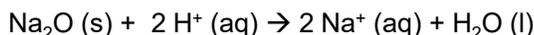
Most follow the pattern **acid + base → salt + water**

It is key to know the charges on the ions so formulae of salts can be deduced e.g. PO_4^{3-} , SO_4^{2-}

The **basic oxides** react with acids to make salts.



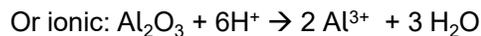
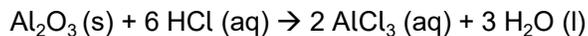
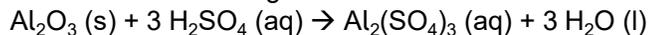
Or ionic equations



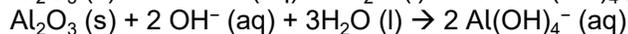
Amphoteric oxides

Aluminium oxide can act as both an acid and an alkali and is therefore called amphoteric.

Aluminium oxide acting as a **base**.



Aluminium oxide acting as a **acid**.



Be careful for whether the question is asking for an ionic equation or a full one

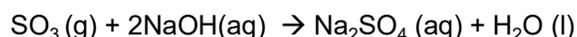
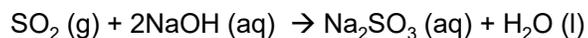
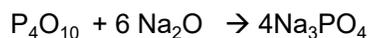
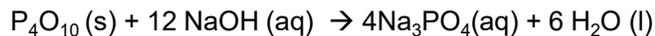
SiO_2 has a **giant covalent structure** with very strong bonds. This stops SiO_2 dissolving or reacting with water and weak solutions of alkali. It will, however, react with very concentrated NaOH



It is still classed as an acidic oxide.

Acidic oxides

The other simple molecular acidic oxides react with bases to form salts.



Or ionic equations:

